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Investigation of the growth parameters of hydrothermal ZnO nanowires for scale up applications

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Abstract

Zinc oxide nano-wires (ZnO NWs) are synthesized reproducibly with high yield via a low temperature hydrothermal technique. The influence of the growth duration time, growth temperature, zinc precursor and base concentration of Na₂CO₃ on the morphology of NWs is investigated. The growth products are characterised using scanning electron microscopy (SEM), Transmission electron microscopy (TEM), X-ray diffraction (XRD), X-ray photoelectron spectroscopy (XPS) and photoluminescence (PL). SEM analysis shows that the optimum growth temperature is 140 °C and finds that length and diameter of ZnO NWs have a relationship with growth duration time and base concentrations of Na₂CO₃. In addition, it is reported that a high (~ 90%) yield of ZnO NWs can be synthesised via using any of three different precursors: zinc chloride, zinc acetate and zinc nitrate. TEM and XRD results indicate the high purity and the single crystalline nature of the ZnO NWs. XPS confirms the absence of sodium contaminants on the surface and indicates a near flat band surface condition. PL shows a large visible band in the yellow part of the spectrum, and a small
exciton emission peak, indicating a large defect concentration, which is reduced after annealing in air.

**Keywords** Semiconducting materials; Nanomaterials; Hydrothermal crystal growth; Characterization; X-ray diffraction; Defects.

1. **Introduction**

Low cost, high yield and controlled fabrication are key factors to realizing the potential of nanomaterials band gap [1, 2], and a high exciton binding energy of 60 meV at room temperature [3-5]. ZnO nanostructures such as nanowires and nanorods are ideal building blocks for many applications due to the unique properties of ZnO, their large surface area to volume ratio [4], and the ability to grow several single crystal geometries.

In recent years, different methods have been used to synthesis ZnO nanowires including vapour liquid solid (VLS) growth [6, 7], metal organic chemical vapour deposition (MOCVD) [7, 8], microemulsion synthesis [9], solvothermal [10] and hydrothermal methods [3]. Amongst all of these methods, hydrothermal growth has attracted a lot of interest because of its many advantages including low cost, low temperature, high yield [3], scalability, and ease of handling. A plethora of nano ZnO shapes have been fabricated using hydrothermal methods including nanobelts, nanowires, nanoflowers and nanoparticles. Hierarchical structures are of particular interest because of their very high surface to volume ratio which has been shown to significantly increase the power conversion efficiency of solar cell [11-13].

ZnO NWs has been grown hydrothermally either in a colloidal form [3, 14] or from a seed layer on various substrates[15-19]. Various studies have found that the growth parameters such as growth duration time [18, 20, 21], growth temperature [22]and precursor concentration [23], have a huge impact on the aspect ratio [3], nanoparticle size [24] and nanowire density [15]. Recent hydrothermal methods have been developed to directly grow NWs at a specific location by localized heat source or using seed patterning. For localized heat, two main process have been used including focused energy field synthesis which relies on Joule heating [25, 26] and laser-induced hydrothermal growth [27-31]. For the seed patterning, methods include microcontact printing of seed nanoparticles [32], inkjet-printed seed patterning [33] and inkjet-printed zinc precursor [34]. These novel methods do not
require expensive photolithography or vacuum deposition and their development widens the range of ZnO NW applications due to their low-temperature synthesis, low cost, environmentally friendly, and simple process setup. Therefore, they can be applied to different types of substrates and can be scaled up for mass production. However, the growth of NWs in a colloidal form is still essential for more versatile applications.

The ZnO NWs studied in this paper are grown using a simple colloidal growth technique first published by Hu et al [3]. In their paper they use a supersaturated solution of Na$_2$CO$_3$ to fabricate NWs over 12 h at 140°C using zinc chloride as precursor. In this paper, we demonstrate the reproducibility of the method and explore the parameter space further, highlighting the high yield of the technique and therefore its potential for scaling up NW fabrication to commercial level. The parameters investigated include the concentration of Na$_2$CO$_3$, growth duration time, the growth temperature and zinc precursors. Additionally, we show that ZnO NWs can be synthesized not only using zinc chloride as precursor salt but also using zinc acetate dihydrate and zinc nitrate hexahydrate. The effect of annealing the grown NWs in air is also studied using characterization techniques not employed in the original paper such as x-ray photoelectron spectroscopy (XPS) and photoluminescence spectroscopy (PL).

2. Experimental procedure

Zinc chloride (ZnCl$_2$), zinc nitrate hexahydrate (Zn(NO$_3$)$_2$·6H$_2$O), zinc acetate dihydrate (Zn(CH$_3$COO)$_2$·2H$_2$O) and sodium carbonate (Na$_2$CO$_3$) were purchased from Sigma Aldrich and used without further purification. The filter paper was purchased from Merck Millipore. In a typical experiment, 0.267 g ZnCl$_2$ (~0.032 M), and 26.67 g Na$_2$CO$_3$ (~4.193 M) were added to a beaker and topped up to 60 mL with deionized water (DI) and stirred at room temperature for 10 minutes. The solution was transferred into a 125 mL Teflon vessel, which was then sealed in an autoclave. The autoclave was placed in an oven and heated to 140 °C for 3 to 26 hours and then cooled down to room temperature naturally. The resulting white products were washed with 8 L of DI water and vacuum filtered using a hydrophilic nylon membrane with a 1.2 µm pore size. Finally, the filtered white products were dried in a desiccator jar for several hours. To obtain a suspension the dried powder sample was sonicated in DI water in an ultrasonic bath for 10 minutes. The experimental procedure using the other precursor salts is identical, keeping the molarity constant at ~0.032 M. The morphology and size of the ZnO products were characterised using a Hitachi model S-4800 II field-emission scanning electron microscope (FE SEM).
measure the length and diameter of NWs several images were taken by SEM at the same magnification (5k for length measurements and 100k for diameter). 100 NWs were chosen randomly from each sample and measured using Image J. Averages, standard deviation and error bars were then calculated (the error bars are 2 standard errors wide). The ZnO NWs crystallinity was examined via a Bruker D8 Discover X-ray diffraction (XRD) system. High resolution Transmission Electron Microscopy (HRTEM) was used to observe the morphology and lattice spacing of the ZnO NWs. The surface chemistry and surface band bending was investigated using a Kratos Supra XPS and the photoluminescence spectra recorded using an Ocean optics USB2000+ spectrometer and a Melles-Griot 325 nm He-Cd laser as the excitation source.

3. Results and discussion

3.1. X-ray diffraction and Transmission electron microscopy results

The crystal structure of ZnO NWs was analysed by powder XRD and TEM. A typical XRD pattern is shown in Fig1. The strongest peaks observed at 2θ values of 31.7, 34.4, 36.2, 47.5, 56.5, 62.8, 66.3, 67.9 and 69.1 are corresponding to the lattice planes of (100), (002), (101), (102), (110), (103), (200), (112) and (201) respectively. All of these diffraction peaks are assigned to a ZnO hexagonal wurtzite structure and are matched with reported data for ZnO (a = 3.249Å, c = 5.205Å JCPDS File, 5-664) [3, 35-37]. Furthermore, no diffraction peaks from Zn or other impurities are observed in the pattern which indicates the high purity of ZnO NWs. High resolution TEM images of a single ZnO NWs are shown in Fig 2. Fig 2(b) shows that the average distance between the lattice planes is 0.26 nm which corresponds to the (0002) plane [38, 39] and therefore a c-axis growth direction.

3.2. Effect of Temperature

The effect of temperature was not reported in the original paper. Here we fixed the concentration of Na₂CO₃ at 4.193 M, the growth time at 6 hours and kept ZnCl₂ as the precursor salt. Separate growth runs were conducted at 120°C, 140°C and 160°C, and the corresponding SEM images of the growth products are shown in Fig 3. Nanowires with diameter of approximately 50 nm and length in excess of 1 µm are only obtained during the 140°C growth. At 120°C, larger crystals with diameter greater than 400 nm and length of 1-2 µm with rough surfaces are produced whilst 160°C yielded predominantly multipod type structures with several large crystals, again with diameter greater than 400 nm, growing out of a common nucleation site. The difference in size of the growth products between 120°C and 140°C is illustrated by the photographs in the insets of Fig3 (a, b) showing that
sedimentation of the suspended growth products 15 mins after suspension is much less at 140°C than at 120°C.

3.3. Effect of growth duration time
NWs were successfully grown at different duration times of 3, 6, 9, 12, 16 and 26 h with all other parameters remaining constant (140°C with 4.193 M Na₂CO₃ and 0.032 M ZnCl₂). Detailed analysis of the SEM measurements is shown in Fig 4(a). The average ZnO NWs length increases with growth duration time from 1.1 µm (3h) to 3.3 µm (26h), whilst the average diameter of the ZnO NWs remains statistically constant at around 41 nm. The length of the NWs obtained after 26 h was not statistically different than the length of the 16 h NWs. This could be explained by the fact that all the free Zn²⁺ in the solution have been consumed. However, the fact that the yield approaches 100% after 6h indicates that the free Zn²⁺ ions supply is already running out after 6 h. Therefore the continuing increase in length with time could be attributed to an Ostwald ripening process where small NWs dissolve and larger NWs keep growing until the process reaches equilibrium after 16 h. The same trend with growth duration time was observed for another set of samples prepared with a lower concentration of 3.145 M Na₂CO₃ and a temperature of 140 °C (not shown). These results suggest that the growth occurs almost exclusively along the c axis once the initial nucleation has occurred. The main reason is likely to be the very high concentration of Na₂CO₃. We have shown that there is a clear link between increasing Na₂CO₃ concentration and decreasing NW diameter. This is possibly a consequence of the alkaline condition induced by the Na₂CO₃ (pH of 11.5). Joo et al [40] have shown that at pH 11, the (1120) side facets of the ZnO NWs become positively charged and the (0001) end facets become negatively charged, resulting in the suppression of lateral growth. It is likely that this mechanism is also responsible for our observations.

3.4. Effect of base concentration (Na₂CO₃)
Na₂CO₃ supersaturation of the solution is an important factor for the nucleation and growth of ZnO NWs [3]. The role of Na₂CO₃ is still unclear but it acts as a weak base which slowly hydrolyses in the water solution and gradually produces OH⁻ [3]. We attempted growth using the following Na₂CO₃ concentrations including 2.044 M, 2.673 M, 3.145 M, 4.193 M and 4.717 M keeping all other parameters constant (140 °C and 6 h). At the lowest concentration of 2.044 M, NWs are not fabricated and only micrometre size crystals are forming, as shown
in Fig 5 (a). When the concentration increases to 2.673 M a mix of NWs and large crystal are obtained as shown in Fig 5 (b). At 3.145 M, 4.193 M and 4.717 M, NWs are successfully fabricated as shown in Fig 5 (c, d and f) respectively. Since the solubility of Na$_2$CO$_3$ at 140 °C is 28.1% by mass [41], the lowest concentration of 2.044 M is below the saturation conditions (21.6% by mass). 2.673 M corresponds to the saturation level with 28.3% by mass and other concentrations including 3.145 M, 4.193 M and 4.717 M are supersaturation solutions with 33.3% by mass, 44.45% by mass and 50% by mass respectively. Based on the above analysis, any concentration below 2.673 M is not able to form ZnO NWs, any concentration above 2.673 M results in supersaturation and hence NWs are synthesized. This is in good agreement with the original paper by Hu et al. However, unlike Hu et al., the formation of nanobelts are not observed. Additionally, our results show that the aspect ratio of the NWs produced is dependent on the base concentration. Fig 4(b) shows that the average length of the NWs increased, whilst the average diameter decreased as the concentration of Na$_2$CO$_3$ increased from 2.673 M to 4.717 M. The length increased from 1.7 µm (2.673 M) up to 2.7 µm (4.717 M), while the diameter decreased from 60.7 nm (2.673 M) down to 40.9 nm (4.717 M). Further addition of Na$_2$CO$_3$ (i.e. above 4.717 M), does not dissolve in the solution and therefore has no effect on the NWs. As mentioned earlier we attribute the decrease in diameter to the increased pH, causing the side facets to become positively charged and suppressing lateral growth.

3.5. Effect of zinc salt precursor
The effect of altering the zinc precursor was also investigated by replacing zinc chloride with zinc nitrate hexahydrate, and zinc acetate dehydrate. All other growth conditions were kept constant: concentration of Na$_2$CO$_3$ at 4.193 M, the growth time at 6 h, the growth temperature at 140 °C and concentration of zinc precursors kept at the same molarity of 0.032 M. The corresponding SEM images of the growth products are shown in Fig 6 (a, b and c). SEM images show that NWs are obtained with all three zinc precursor used. Analysis of the SEM measurement is shown in Fig 6. It can be seen that NWs have approximately the same length and diameter which means that the zinc salt precursors do not have a significant impact on the aspect ratio or morphology of NWs.

3.6. ZnO nanowires yield
Fig 5(f) shows a digital photograph of the nanowires lifted off the filtration membrane after washing, filtration and drying in a desiccator. The filtered products for three identical growth runs (3.145 M Na$_2$CO$_3$, 0.032 M ZnCl$_2$, 140 °C and 6 h) were weighed and an average
weight of 142.8 mg of ZnO nanowires was obtained, which correspond to an atomic yield of nearly 90%; i.e. 9 out of 10 zinc atoms provided by the precursor were used to produce ZnO nanowires. This remarkable high yield, associated with the low cost of the chemicals and the ease of fabrication, makes this method very favourable for scale up operations.

3.7. X-ray photoelectron spectroscopy results
The XPS survey scan in Fig 7 showed the expected Zn, O and C peaks and demonstrates the purity of the sample after rinsing and filtering as no Na peak is present in spite of the high concentration of Na₂CO₃ during the growth. Percentage composition of the O1s core level, Zn 2p and C 1s as a function of annealing temperatures in air are shown in table 1, extracted from the curve fitting of each core level. The O1s core level peak shape is shown in Fig 8 and fitted with two components. The first peak at 530.7 eV is attributed to O²⁻ oxidation state bound with Zn in the crystal lattice and the high binding energy peak at 532 eV corresponds to surface oxygen species such as OH⁻ ions[42-45]. Table 1 indicates that there is a significant reduction in surface carbon and surface oxygen with temperature. The zinc to lattice oxygen ratio decreased from 1.145 to almost 1 following the last anneal indicating that annealing resulted in a reduction of oxygen vacancies. In addition, XPS valence band spectra shown on Fig 9 reveal a near flat band condition with the valence band edge 3.1eV below the Fermi level, which was unaffected by the anneals.

3.8. Photoluminescence results
The PL spectra of NWs grown for 3, 6, 9, and 12 h are shown in Fig 10, as well as the spectrum of the 3h NWs after annealing at 400 °C. The spectra were normalized to the excitonic, near band edge (NBE) peak at 376 nm in the main graph and to the deep level emission (DLE) band in the inset. The DLE band is thought to originate from several radiative transitions between the band edges and deep levels caused by oxygen vacancies and interstitials[46-50]. The main graph shows that the DLE peaks increase in intensity compared to the NBE peak with growth duration time but there is no change in the shape of either peak. This suggests that the efficiency of radiative recombination through deep levels appears to increase with growth duration time, compared to exciton recombination. The shape of the
DLE band, with an intensity maximum near 600nm is also very similar to previously reported results on vertical ZnO NWs arrays [51]. Annealing the 3h sample at 400°C (inset of Fig 10) increases the intensity of the NBE peak relative to the DLE band by a factor of 4. The anneal also causes the DLE band shape to change, with the peak maximum shifting from 600 nm to 640 nm, indicating that the lowest wavelength defect assisted transition is decreased by the anneal compared to the other contributions to the DLE. The shape of the DLE band after annealing at 400°C is also very similar to that observed on ZnO nanosheets produced by annealing zinc acetate at 400°C in air [47, 52], indicating that the effect of this anneal are reproducible on different ZnO nanomaterials. Both the increase in NBE emission and the change in shape of the DLE band were also observed on the 12h sample following annealing at 400°C (not shown). It can be concluded that the anneal removes defects, improving crystallinity. This is in good agreement with the XPS showing a reduction in oxygen vacancies.

4. Conclusion
The growth parameters (temperature, duration time and base concentration) of a low temperature hydrothermal method for producing ZnO NWs with controllable length and diameter have been investigated. The optimum growth temperature was found to be centred around 140°C. Increasing the growth duration time increases the length of the grown ZnO NWs but does not have a significant impact on the NWs diameter. Increasing the base concentration increases the aspect ratio of the NWs. Three different salt precursors were used with no significant difference in the growth products. The yield was measured at around 90%. XPS investigation shows effective removal of surface contamination of ZnO NWs annealed in air up to 400°C and a reduction in oxygen lattice vacancies. XPS also indicated near flat band conditions at the surface, unaffected by air annealing. Finally, PL shows that the 400°C anneal removes defects, improving the intrinsic optical properties of the material. This method is simple, low cost, requires no harmful chemicals and possesses a very high yield and so could offer a viable route for the scaled up production of ZnO NWs with controllable aspect ratios.

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References


[18] Han Z, Li S, Chu J, Chen Y 2013 Controlled growth of well-aligned ZnO nanowire arrays using the improved hydrothermal method J. Semicond. 34 063002


[31] Yeo, J, et al, 2015 Laser-induced hydrothermal growth of heterogeneous metal-oxide nanowire on flexible substrate by laser absorption layer design ACS nano 9 6059-68
[34] Kwon, J, et al, 2013 Direct selective growth of ZnO nanowire arrays from inkjet-printed zinc acetate precursor on a heated substrate Nanoscale research letters 8 489
[38] Kamble AS et al 2014 Effect of hydroxide anion generating agents on growth and properties of ZnO nanorod arrays Electrochim. Acta 149 386-93


[46] Djurišić A et al. 2007 Defect emissions in ZnO nanostructures. *Nanotechnology.* **18** 095702


[52] Tarat A et al 2014 Microwave-assisted synthesis of layered basic zinc acetate nanosheets and their thermal decomposition into nanocrystalline ZnO *Nanoscale Res Lett.* **9** 11
**Table 1**: Percentage composition of the O1s core level, Zn and C 1s as a function of annealing temperatures in air

<table>
<thead>
<tr>
<th>Annealing Temperature °C</th>
<th>Oxygen lattice</th>
<th>Oxygen surface</th>
<th>Zinc</th>
<th>Carbon</th>
<th>Zinc/lattice Oxygen</th>
</tr>
</thead>
<tbody>
<tr>
<td>70</td>
<td>20.54</td>
<td>25.99</td>
<td>23.52</td>
<td>29.95</td>
<td>1.145</td>
</tr>
<tr>
<td>200</td>
<td>24.90</td>
<td>21.55</td>
<td>26.85</td>
<td>26.7</td>
<td>1.078</td>
</tr>
<tr>
<td>300</td>
<td>28.84</td>
<td>19.92</td>
<td>29.71</td>
<td>21.54</td>
<td>1.030</td>
</tr>
<tr>
<td>400</td>
<td>31.38</td>
<td>20.19</td>
<td>31.50</td>
<td>16.92</td>
<td>1.003</td>
</tr>
</tbody>
</table>
Fig. 1 XRD patterns recorded from the ZnO nanowires.

Fig. 2. (a) HRTEM image of a ZnO NW. The inset shows the location of the image acquisition. (b) High magnification HRTEM image taken from the area indicated in (a) and showing the distance between the (0002) planes.
Fig. 3 ZnO NWs grown at (a) 120 °C, (b) 140 °C and (c) 160 °C. All other parameters are identical and the three images are recorded at the same magnification. The insets in (a) and (b) show suspension of the rinsed growth products 15 mins after sonication, highlighting the difference in sedimentation rate.

Fig. 4(a) ZnO NWs length and diameter versus growth duration time. NWs are grown at 140 °C using a base concentration of 4.193 M. 4(b) ZnO NWs length and diameter versus base (Na₂CO₃) concentration. NWs are grown at the same temperature of 140 °C and for 6 h. Error bars are 2 standard errors wide.
Fig. 5 SEM Images of the ZnO NWs grown at 140 °C for 6 h with base concentrations of Na$_2$CO$_3$ of (a) 2.044 M, (b) 2.673 M, (c) 3.145 M, (d) 4.193 and (e) 4.717 M. Images are taken at the same magnification. (g) Filtered and dried ZnO NWs produced from 0.032 M ZnCl$_2$.

Fig. 6 ZnO NWs SEM images, length and diameter for the three zinc precursors studied. Growth conditions are the same apart from the zinc salt: 140°C, 6 h, 4.193 M Na$_2$CO$_3$ and 0.032 M zinc salt.
Fig. 7 XPS survey scan of the ZnO NWs

Fig. 8 XPS spectra of the O 1s core level from the as grown NWs. The raw data was fitted with two components and a Shirley background. The two components correspond to the oxygen crystal lattice and the surface oxygen.
3.8. Photoluminescence results

Fig. 9 XPS valence band spectra of ZnO NWs as grown, and annealed at 200, 300 and 400 °C.

Fig. 10 PL of ZnO NWs grown for 3, 6, 9 and 12 h, and NWs grown for 3 h and subsequently annealed at 400°C. The spectra are normalized to the near band edge peak at 376 nm. The inset shows the spectra of the 3 h NWs, before and after annealing, and are normalised to the deep level band (DLE) intensity to highlight the change in the shape of the DLE band.